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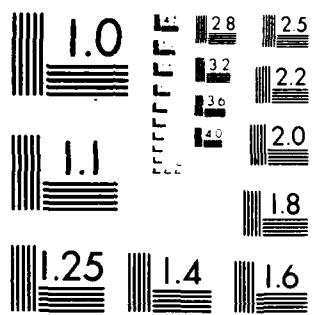
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SPECTROSCOPIC PROPERTIES OF METAL MONOXIDES
AND HYDROXIDES IMPORTANT IN THE IONOSPHERE

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e/f -dependence of $B \leftarrow a$ intensity factors enables, for the first time, measurement of relative $a^3\Pi$ and $A^1\Pi$ populations from observed $B \rightarrow X$ fluorescence intensities. A significant excess population (relative to $A^1\Pi$), above the expected 500K thermal distribution is found in the e-parity levels of $a^3\Pi$ ($v_a = 0$ and 1) in a 1-2 Torr Mg + Ar + N₂O Broida-oven flame. Whether these e vs. f and $a^3\Pi$ vs. $A^1\Pi$ population anomalies reflect the nascent product distribution or result from complex, multiple collision processes is under investigation.

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The most important use of the MgO/B-a system will be proof, via OODR experiments now in progress, that the $\lambda < 320\text{nm}$ continuum arises from the MgO $e^3\Sigma^-$ ($\text{Mg}^{+1S} + \text{O}^{+3P}$) $\rightarrow a^3\Pi$ transition, thereby providing direct optical spectroscopic bounds to the dissociation energy of the MgO molecules.

Four papers are in preparation describing the results of this research program.

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SPECTROSCOPIC PROPERTIES OF METAL MONOXIDES AND
HYDROXIDES IMPORTANT IN THE IONOSPHERE

A. INTRODUCTION

The monoxides, monoxide ions, and hydroxides of Mg, Ca, and Fe are important species in the ionosphere. Spectroscopic information about these species, needed for atmospheric modelling, is either incompletely available or the subject of controversies. This program of research was originally planned to characterize the lowest lying states of MgO, CaO, and FeO and to provide a direct spectroscopic measure of the dissociation energy of MgO.

Although several projects dealing with the low lying states and electronic structure of CaO^{1,2,3} and CuO⁴ were completed during the period of this contract, the present report will deal exclusively with the MgO molecule.

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B. SUMMARY OF RESULTS

1. MgO B¹Σ⁺-a³Π_J System

The a³Π_J state is the first excited state of MgO. It was first identified through its perturbations of the X¹Σ⁺ state.⁵ Although the a³Π_J state is an important "reservoir state" in Mg oxidation reactions and the e³Σ⁻-a³Π_J system has been suggested as responsible for a $\lambda < 320\text{nm}$ continuum (forming Mg ¹S + O ³P atoms)⁶, no suitable scheme exists for monitoring populations in specific MgO a³Π rotation-vibration levels.

The main thrust of this program was to record, assign, and deperturb the MgO B¹Σ⁺-a³Π_J intercombination system. Although nominally forbidden, the B-a system borrows oscillator strength from the B¹Σ-A¹Π and B¹Σ-X¹Σ systems through spin-orbit interactions of the a³Π state with A¹Π and X¹Σ⁺. The net result of the A-a mixing is that the B¹Σ⁺-a³Π($\Omega = 1$) sub-band is computed to have $\sim 1/450$ of the oscillator strength of B-A bands. However, a surprisingly large excess population in the a³Π e-parity levels has made recording parts of the B-a system much easier than initially expected. Unfortunately, the extremely small populations in f-parity levels and the weakness of transitions out of $\omega = 0$ and 2 components have made characterization of the MgO a³Π state unexpectedly difficult.

Table I shows the J = 0 - 50 lines of the complete B - a (0,0) band. Two complementary techniques were used: cw and pulsed dye laser fluorescence ($B^1\Sigma^+ \rightarrow X^1\Sigma^+$, $\Delta V = 0$) excitation spectroscopy (FES). The lower intensity cw laser only weakly saturates the B-A system, thus the relative B-A/B-a intensities sample populations times line strength factors. The high intensity

TABLE I
MgO B₁ + a₁ (0,0) Observed and Calculated Line Positions (cm⁻¹)^a

J	9-8 ^{b,c}			B-a ^{b,c}			B-a ^{b,c}			B-a ^{b,c}		
	P(J)	Q(J)	R(J)	P(J)	Q(J)	R(J)	P(J)	Q(J)	R(J)	P(J)	Q(J)	R(J)
0	17381.377C	17386.356C	17383.553C	17452.301C	17453.457(-7)	17455.782C	-	-	-	17512.502C	17514.822C	17518.302C
1	380.508C	386.500C	384.858C	451.466(1)b	454.785(1)	457.257(2)	458.892(8)	511.870C	515.350C	519.990C	515.350C	519.990C
2	379.783C	386.788C	387.908C	450.769(5)	454.769(5)	457.257(2)	458.892(8)	511.870C	515.350C	521.853C	516.053C	521.853C
3	379.206(4)d	387.795C	389.631(-12)	450.230C	454.898(16)	450.669(-1) b	450.669(-1) b	511.133C	516.933C	523.893C	516.933C	523.893C
4	378.766(-2)	388.515C	391.528C	449.849(-3)	455.695(9)	462.589(-23) b	462.589(-23) b	511.712(2)	517.989C	526.198C	517.989C	526.198C
5	378.485(7)b	389.379C	393.619(12)b	449.633C	456.388C	467.752(3)	466.964(-1)	511.104C	519.220C	528.499C	519.220C	528.499C
6	378.328(-6)b	390.386C	395.739(7)b	449.573(6)	459.573(6)	467.752(3)	466.964(-1)	511.104C	519.220C	528.499C	519.220C	528.499C
7	378.328(-7)b	391.539C	398.031(-2)d	449.658(-1)	459.021(2)	467.372(-4)	467.372(-4)	511.388C	520.627C	531.064C	520.627C	531.064C
8	378.485(3)b	392.835C	400.516C	449.910(1)	450.444(-3)	471.945(2)	471.945(2)	511.771C	522.309C	533.805C	522.309C	533.805C
9	378.776(11)	394.276C	403.112(11)	450.312(-4) b	454.656(-11) b	471.656(-11) b	471.656(-11) b	512.389C	523.721C	534.721C	523.721C	534.721C
10	379.206(-7)b	395.862C	405.880(2)	450.874(-5)	463.800(9) b	477.548C	477.548C	513.144C	525.900C	539.813C	525.900C	539.813C
11	379.800(3)	397.592C	408.784(2)	451.594(-6)	465.683(-3)	480.58(-1)	480.58(-1)	514.593C	526.309C	543.078C	526.309C	543.078C
12	380.521(-5)	399.467C	411.827C	452.481(3)	467.749(1)b	483.784(5)b	483.784(5)b	515.282C	529.220C	546.512C	529.220C	546.512C
13	381.597(-6)	401.488C	415.056(-3)	452.542(-2)	465.462(-1)d	481.122(-1) d	481.122(-1) d	516.517C	522.156C	556.132C	522.156C	556.132C
14	382.432(6)	403.655(2)	418.358(3)	454.714(6)	472.348(-7) b	490.639(2) b	490.639(2) b	517.992C	535.382C	553.923C	535.382C	553.923C
15	383.520(-4)	405.973(8)	421.839(2)	456.068(9)	474.900(4)	494.301(-1)	494.301(-1)	519.641C	536.887(-3)	557.875(-8)	536.887(-3)	557.875(-8)
16	384.905(-6)	408.416(-5)	411.018(-6)	429.235(-2)	457.234(-1) d	476.601(5)	476.601(5)	517.460(-4) b	533.455(-2)	562.014(-7)	533.455(-2)	562.014(-7)
17	386.373C	411.721(10)	433.491(-10) b	461.050(-1) b	480.455(-1) b	502.101(1)	502.101(1)	525.621(-1) d	544.655(-1)	566.328(-8)	544.655(-1)	566.328(-8)
18	387.912(10)	413.721(10)	433.721(10)	463.336(-7)	483.473(4)	506.237(1)	506.237(1)	525.621(-1) d	544.655(-1)	570.805C	544.655(-1)	570.805C
19	389.721(-2)d	416.664(4)	437.226(2)	463.336(-7)	486.652(1)	509.529(1)	509.529(1)	527.994(4)	551.149(-10)	575.466(-2)	551.149(-10)	575.466(-2)
20	391.648(5)	419.181(9)	441.440(4)	465.185(1)	489.971(2) d	510.510(1)	510.510(1)	530.151(3)	550.841(5)	580.294C	550.841(5)	580.294C
21	393.695(1)	422.898C	445.801(6)	467.481(-4) b	493.486(1)	519.588(13)	519.588(13)	533.190(3)	558.624(-12)	585.294C	558.624(-12)	585.294C
22	395.896(3)	426.227(-6)	450.312(13)b	469.96(2)b	497.139(-2)	524.366(18) b	524.366(18) b	536.606C	562.115(1)	59.466C	562.115(1)	59.466C
23	398.243(3)	429.715C	454.952(1)	472.560(-2)	500.957(1)	529.270(2)	529.270(2)	539.981(1)	559.907(3)	595.809C	559.809C	595.809C
24	400.735C	433.347(2)	459.535(1)	475.535(-1) d	504.557(5)	532.155(4)	532.155(4)	551.557(5)	571.529(9)	601.326C	571.529(9)	601.326C
25	403.381(2)	437.123(1)	464.694C	484.274(1)	509.060(6)	539.594(5)	539.594(5)	545.066C	575.621(8)	602.011C	575.621(8)	602.011C
26	406.171(1)	441.049(1)	469.786(-1) d	481.367(-2)	511.35(-3)	544.984(1) b	544.984(1) b	549.233C	580.526C	612.886C	580.526C	612.886C
27	409.114(3)	445.122C	475.619(-1) b	484.823C	517.83C	542.936(1)	542.936(1)	552.981(5)	586.411C	618.89C	586.411C	618.89C
28	412.199(-2)	449.346(1)	480.411C	488.032(-4)	522.420(-7)	556.249(2)	556.249(2)	566.386(2)	596.466C	625.395C	596.466C	625.395C
29	415.441(2)	453.716(-1)	485.34(-1) d	491.616(7) b	522.191(-1) b	552.123(8)	552.123(8)	566.231(7)	595.994C	631.463C	595.994C	631.463C
30	418.836(9)	458.240(3)	491.616(-1) b	495.340(-2)	532.134C	568.441C	568.441C	586.231(1)	601.394C	638.001C	601.394C	638.001C
31	422.365C	462.911(4)	497.231(-3)	503.285(-1)	542.486C	573.213(-9)	573.213(-9)	594.236C	614.664C	644.709C	614.664C	644.709C
32	426.053(1)	467.727C	503.424(-10)	542.417(3)	578.362C	605.362C	605.362C	624.214C	651.495C	681.586C	651.495C	681.586C
33	429.889C	472.692(-4)	509.561(1)	507.492(-6)	545.499(6)	583.737C	583.737C	611.186C	635.355C	665.835C	635.355C	665.835C
34	433.876C	477.813(2)	515.837(4)	511.871(1)	553.466C	593.226C	593.226C	612.226C	642.455C	672.664C	642.455C	672.664C
35	439.014C	483.554(-2)	528.621(4)	545.40(-1)	595.200C	636.643C	636.643C	656.643C	686.643C	716.735C	686.643C	716.735C
36	446.008C	485.481(-3) b	528.621(4)	545.095(2)	595.949C	636.600(-3)	636.600(-3)	656.600(-3)	686.600(-3)	721.345C	686.600(-3)	721.345C
37	446.365C	494.579C	535.548(2)	525.945C	572.144C	614.49C	614.49C	634.49C	664.765C	701.598C	664.765C	701.598C
38	446.744C	496.727C	536.705(4)	528.99C	578.407(7)	615.842C	615.842C	635.842C	665.842C	705.311C	665.842C	705.311C
39	451.331C	499.303(1)	540.542(1)	534.417(3)	580.550(3)	616.842C	616.842C	636.842C	666.842C	706.386C	666.386C	706.386C
40	456.072C	496.411C	549.581(13)	536.127(-1) b	586.127(-1) b	618.487C	618.487C	638.487C	668.487C	707.420C	668.487C	707.420C
41	460.964C	511.699(-6)	556.597C	541.457(1)	596.273C	627.936C	627.936C	647.936C	678.936C	712.664C	678.936C	712.664C
42	466.008C	512.885(1)	563.912C	546.950(-2)	596.956C	634.955C	634.955C	654.955C	684.955C	721.141C	684.955C	721.141C
43	471.203C	524.216C	557.371(-2)	540.600(-3)	597.825C	632.766C	632.766C	652.766C	682.766C	721.723C	682.766C	721.723C
44	476.549C	530.705(4)	578.99C	558.407(-7)	601.842C	635.76C	635.76C	655.76C	685.76C	725.311C	685.76C	725.311C
45	482.148C	537.345(1)	586.753C	564.184C	603.020C	639.084C	639.084C	659.084C	694.095C	734.204C	694.095C	734.204C
46	487.698C	551.072C	596.644C	534.457(1)	576.803C	636.457C	636.457C	656.457C	696.457C	736.455C	696.457C	736.455C
47	493.500C	558.175(3)	610.937C	583.249C	640.518C	674.732C	674.732C	694.732C	734.822C	774.867C	734.822C	774.867C
48	499.454C	565.425C	619.497C	589.555C	648.339C	673.592C	673.592C	693.592C	730.620C	771.312C	730.620C	771.312C
49	505.560C	572.832C	627.806C	596.617C	656.322C	692.405C	692.405C	712.405C	750.426C	784.396C	750.426C	784.396C
50	511.812C	-	-	-	-	-	-	-	-	-	-	-

^a Number in parentheses is (observed - calc.) in 10⁻³ cm⁻¹.^b Blended lines.^c Calculated lines.^d These lines are obtained by using a pressure-scanned pulsed dye laser with a 44MM of 0.02cm⁻¹. Experimental uncertainties on these lines are 0.04cm⁻¹.

TABLE II

Constants for the a, B, and X States of MgO (cm^{-1})^a

	a ^{3,1}	B ^{1,2}	X ^{1,2}			
	v = 0	v = 1 ^b	v = 0	v = 0	v = 3	v = 4
τ_0	2553.755(19)	3195.4625(16)	20003.575(11)	0.0d	2293.67(43)	3038.54(46)
B_V	0.500341(12)	0.495556(7)	0.580058(12)	0.571987(25)	0.5568c	0.55151(19)
$D_V \times 10^6$	1.1979(49)	1.1072(40)	1.1485(47)	1.2173(97)	1.29c	1.31c
A_V	-63.489(14)	-64.1727(73)	-	-	-	-
$A_{D_V} \times 10^4$	-1.214(27)	-2.953(80)	-	-	-	-
A_{Q1}	17.69(27)	20.889(29)	-	-	-	-
A_{11}	-65.11(11)	-64.85c	-	-	-	-
O_V	-1.046(28)	-0.619c	-	-	-	-
P_V	-0.02001(25)	0.02391(36)	-	-	-	-
$Q_V \times 10^4$	-0.96(10)	-1.96c	-	-	-	-
$O_{D_V} \times 10^4$	-0.741(88)	-	-	-	-	-

^a Number in parentheses is one standard deviation uncertainty.^b Preliminary values. Constants for B^{1,2} were fixed.^c Held fixed in the fit.^d Zero of energy.

pulsed laser saturates both transitions, hence sampling relative populations. Although B-a lines are weaker than B-A in the cw laser spectra, the B-A lines are barely detectable in the pulsed laser spectra! The use of pulsed laser FES will provide the crucial missing $\Omega = 0$ and 2 lines in the B-a (0,1) band.

Table II gives the constants for the $a^3\Pi v = 0$ and 1 levels obtained from the measured lines of the B-a (0,0) and (0,1) bands. Reduction to a set of mechanically and magnetically meaningful constants required deperturbation of $a^3\Pi v_a$ with respect to interactions with $X^1\Sigma^+ v_X = v_a + 3$ and $A^1\Pi v_A = v_a$. In turn this deperturbation model yields precise basis function mixing coefficients which enable calculation of v, J, Ω , and e/f dependent B-a intensity factors.

2. Intensities in the MgO $B^1\Sigma^+$ - $a^3\Pi_g$ System

In order to make quantitative measurements of the relative populations of $a^3\Pi_Q e$ vs. f , $a^3\Pi_Q$ vs $a^3\Pi_{Q+1}$, or $a^3\Pi_Q$ vs $A^1\Pi$ levels, it is necessary to know the relative linestrength factors. Such factors are never trivially available for a nominally forbidden band system. Fortunately, the $B + a$ (v_B, v_a) band of the FES spectrum monitored by $B \rightarrow X$ ($v_B, v_X = v_B$) fluorescence, borrows more than 99% of its intensity from the $B-A$ ($v_B, v_A = v_a$) band. This means that knowledge of the $A^1\Pi$ character in each nominal $a^3\Pi v_a, J, \Omega, e/f$ level is sufficient to calculate the intensity of any $B-a$ line relative to any $B-A$ or other $B-a$ line. The $A^1\Pi$ fractional character varies over several orders of magnitude, and the mixing fractions obtained from the deperturbation model are accurate to better than 1% of their calculated magnitudes. The deperturbation model yields mixing fractions which should be considerably more accurate than state-of-the-art absolute intensity measurements.

Although unimportant for most purposes, quantum mechanical interferences between $B-A$ and $B-X$ transition amplitudes can cause certain $B-a$ lines to behave anomalously. P and R lines from a common J'' level would normally have comparable linestrengths. However, for $B + a$ bands detected by $B \rightarrow X \Delta v \neq 0$ fluorescence, the P/R intensity ratio will deviate significantly from 1. Whether it is greater or less than 1 determines the relative signs of the $B - X$ and $B - A$ transition moments.

For the first time, populations in the most important "reservoir state" of MgO, the MgO $a^3\Pi$ state, can be monitored systematically and without assignment ambiguity. The monitoring transition is in the convenient Rhodamine 6-G region and had the simplest possible branch structure. Although

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the d-a system has been used for monitoring $a^3\Pi$, its complex structure has, to date, frustrated analysis. The availability of $a^3\Pi$ combination differences should facilitate assignment of the $d^3\Delta$ - $a^3\Pi$ system and resolution of several contradictory features of speculative bandhead assignments.^{5,7}

3. Relative $a^3\Pi/e/a^3\Pi/f$ and $a^3\Pi/A^1\Pi$ Populations.

Armed with the calculated linestrength factors from the $a^3\Pi \sim A^1\Pi, X^1\Sigma^+$ deperturbation, it becomes possible to measure population ratios and to compare these ratios to those expected from a thermal ($T \approx 500K$) distribution. At 1-2 Torr total pressure (99% Ar, 1% N₂O, 0.1% MgO and Mg), it is reasonable to expect the metastable X, A, and a levels to be equilibrated. However, the $a^3\Pi v = 0, \varpi = 1$ levels have more than a factor of 10 higher population relative to $A^1\Pi v = 0$ than the 950 cm⁻¹ separation of these levels would imply at 500K. Similarly, the $a^3\Pi \varpi = 1$ e-levels have more than a factor of 2 greater population than the near degenerate ($\Delta E < 0.1$ cm⁻¹) f-levels. It is difficult to imagine how such large departures from equilibrium could exist at such high pressure.

Single collision Mg +N₂O experiments, in collaboration with Professor Paul Dagdigian at Johns Hopkins University, are in progress. The purpose of these experiments is to determine whether the 1-2 Torr population anomalies reflect nascent product distributions or complex multiple-collision population funnelling and removal processes.

4. Dissociation Energy of MgO

Ground state Mg and O atoms correlate adiabatically with the MgO $a^3\pi$ and $e^3\Sigma^-$ states. The e-a transition is predicted to be strong because the $e^3\Sigma^-$ state belongs to the same electronic configuration as d^3A_1 , D^3D , and $C^3\Sigma^-$. Evans and Mackie^b have observed a continuum in absorption from shock-heated MgO at $\lambda < 320$ nm which they have assigned as $e^3\Sigma^- + a^3\pi$. If this assignment can be verified and the long- λ limit of the continuum can be located with greater sensitivity, it should be possible to obtain a direct spectroscopic upper bound to D_0^∞ (MgO).

Our initial purpose for investigating the MgO B-a system was to develop and unambiguous $a^3\pi$ population monitor for use in an Optical-Optical Double Resonance (OODR) scheme for assigning and extending the Evans-Mackie continuum. Laser 1 destroys MgO $a^3\pi$ by pumping in the continuum. Laser 2 detects MgO $a^3\pi$ and shows that the population destruction which occurs when Laser 1 pumps in the continuum is specific to $a^3\pi v = 0$. If the Evans-Mackie assignment is correct, then population loss in $A^1\pi$ (monitored by Laser 2 in the B-A system) will occur at a longer delay after the Laser 1 pulse than population loss in $a^3\pi$.

The pulsed laser FES component of this experiment has been demonstrated. The full OODR experiment will be performed in March, 1983.

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PUBLICATIONS

The following papers are in preparation:

1. Precila C.F. Ip, Robert W. Field, and Keith J. Cross, "The $B^1\Sigma^+ - a^3\Pi_j$ Intercombination System of the MgO Molecule," for J. Mol. Spectrosc.
2. Keith J. Cross and Precila C.F. Ip, "The MgO $B^1\Sigma^+ - A^1\Pi$ System. Laser Spectroscopy and Deperturbation," for J. Mol. Spectrosc.
3. Precila C.F. Ip and Robert W. Field, "Intensity Factors for the MgO $B^1\Sigma^+ - a^3\Pi_j$ System Based on $X^1\Sigma^+ - a^3\Pi_j - A^1\Pi$ Deperturbation Model," for Chem. Phys. Lett.
4. Precila C.F. Ip, Keith J. Cross, and R.W. Field, "Population Anomalies in an Mg+N₂O Flame. Excess $a^3\Pi/A^1\Pi$ and $a^3\Pi_e/a^3\Pi_f$ Populations," for J. Phys. Chem.

INTERACTIONS AND PRESENTATIONS

- 6-80 Molecular Spectroscopy Symposium, Columbus, Ohio, P.F. Bernath, P. Ip, and R.W. Field, "Optical-Optical Double Resonance Spectroscopy of BaF" (contributed talk MF5).
- 3-81 Yale University, Symposium on Lasers in Chemistry, R.W. Field, "Molecular Spectroscopy Beyond Molecular Constants" (invited).
- 4-81 Discussion of Faraday Society, No. 71, Bristol, England, R.W. Field, "Tunable Laser Electronic Spectroscopy" (invited).
- 5-81 International Colloquium on Molecular Spectroscopy, Stockholm, Sweden, R.W. Field, "Perturbations: Sublime and Ridiculous" (invited).
- 6-81 Molecular Spectroscopy Symposium, Columbus, Ohio, K.J. Cross, P. Ip, and R.W. Field, "Rotational Analysis of the $B^1\Sigma^+ - a^3\Pi_g$ Band System of MgO (contributed talk MG7).
- 3-82 Bad Honnef, Germany, conference of the Deutsche Bunsengesellschaft für Physikalische Chemie on Small Molecules in the Gas Phase, R.W. Field, "Electronic Structure of Diatomic Molecules Beyond Simple Molecular Constants" (invited).
- 8-82 Gordon Conference on Molecular Electronic Spectroscopy, R.W. Field, "The Electronic Structure of Ionic Diatomic Molecules" (invited).

Collaborations with Dr. Guy Taieb and Dr. Bernard Bourguignon at Université de Paris-Sud, Orsay and Professors Yarkony and Dagdigian at Johns Hopkins University.